Support for an anaerobic sulfur cycle in two Canadian peatland soils

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1 Sulfur cycling in peatlands may affect global CH4 emissions by suppression of methanogenesis through bacterial sulfate reduction (BSR). We sought evidence for anaerobic sulfur cycling in four peat mesocosms irrigated with sulfate at a loading of 0.8 and 3.3 g S m−2 yr−1. To this end we obtained concentration profiles of dissolved O2, C, S, and Fe, and determined 34S/32S ratios of sulfate, reduced inorganic sulfur (TRIS), and total sulfur. To estimate the importance of BSR for anaerobic respiration, peat was incubated with molybdate as inhibitor of BSR. In the mesocosms, pore water concentrations of dissolved sulfate and H2S adjusted to 5–20 μmol L−1 and 0–9 μmol L−1, respectively, whereas concentrations of CO2, CH4, and DOC reached millimolar levels. CO2 production was not explained by methanogenesis and net reduction of inorganic electron acceptors. In the shallow peat, H2S was produced and 34S in sulfate enriched by 3.6 to 6%, indicating occurrence of BSR. Sulfate reducers also accounted for much of the metabolic activity. Addition of molybdate suppressed CO2 production by 20 to 50%. Deeper into the peat, the sulfate pool was apparently replenished from the peat matrix as sulfate became enriched in 32S, likely stemming from TRIS or organic sulfur in the peat. Sulfur was thus anaerobically cycled between oxidized and reduced pools. An electron acceptor capable of driving this cycle could not be conclusively identified. Regardless of this uncertainty, the results suggest that anaerobic S cycling can maintain BSR and potentially contribute to low methane production in soils of ombrotrophic bogs.


1. Introduction

[2] The fate of sulfur (S) deposited on the surface of wetlands and peatlands and its effects on carbon (C) cycling has seen a resurgence in interest, owing to the possible ramifications for greenhouse gas emissions from such environments [Gauci et al., 2002; Vile et al., 2003b]. Sulfate is important for rates and pathways of anaerobic carbon mineralization because its presence may decrease methane production through competition of anaerobic bacteria for substrates, such as acetate, alcohols, and hydrogen [Conrad, 1999; Kotsyurbenko et al., 2004; Shannon and White, 1996]. Rates of bacterial sulfate reduction (BSR) have been reported to exceed rates of methane production in freshwater peats, on the basis of samples of peat soil incubated in vitro with tracer amounts of 35S-sulfate [Nedwell and Watson, 1995; Vile et al., 2003a, 2003b; Wieder et al., 1990]. These findings have been used to conclude that sulfate deposition inhibits methanogenesis through competitive suppression of methanogens by sulfate reducers [Gauci et al., 2002; Nedwell and Watson, 1995]. A net decrease of cumulative emissions of atmospheric methane has been verified in sulfate deposition experiments [Dise and Verry, 2001; Gauci et al., 2002].

[3] This hypothesis is not without caveats, though. The sum of methanogenesis and BSR often only explains a small fraction of anaerobic CO2 production in ombrotrophic peats [Vile et al., 2003a]. Electron accepting processes other than BSR may, therefore, be more important for anaerobic CO2 production, which raises the question why a large effect of sulfate deposition on methanogenesis should be expected. Uptake of sulfate by plants, most of which was absorption by Sphagnum mosses, was furthermore reported to be about 0.9 g S m−2 yr−1 in a central continental bog, which exceeds moderate atmospheric sulfate deposition rates [Urban et al., 1989]. Active sulfate reduction, which has still been reported from similar peatlands, should thus be primarily fueled by mineralization of organic S in the peat rather than sulfate deposition. Further problems arise from mass balance considerations. Sulfate pools are small in ombrotrophic peatlands and turn over rapidly [Nedwell and Watson, 1995; Wieder et al., 1990]. The only established mechanism to allow for S recycling in saturated environments poor in ferric iron hydroxide involves reoxidation of reduced S with oxygen in the rhizosphere of plants capable of aerenchymatic transport [Wind and Conrad, 1997].
mechanism, however, is of little significance in Sphagnum and shrub dominated continental bogs. We clearly need to gain better insight into the fate of sulfate in peatlands, especially with respect to our understanding of CH$_4$ production and emission.

Several processes have to be considered regarding the fate of S in the saturated zone of organic rich soils (Figure 1). Sulfate may form organic sulfate esters (SO$_4^2$-R) or be reduced by sulfate reducing bacteria (SRB). Hydrogen sulfide may subsequently be chemically incorporated into reduced inorganic sulfur (TRIS) or carbon-bonded sulfur (CBS), or be chemically or microbially reoxidized to thiosulfate (S$_2$O$_3^{2-}$), for example by quinone moieties (DOM-Q) contained in dissolved organic matter that leaches from vegetation and the unsaturated zone. Sulfate reduction leaves the remaining sulfate enriched in $^{34}$S and the produced H$_2$S, TRIS, and CBS depleted in $^{34}$S along the flow path (strong isotope fractionation [Canfield, 2001a]). Replenishment of the sulfate pool from H$_2$S, TRIS, and CBS will reduce the $^{34}$S content of sulfate along the flow path (small isotope fractionation [Canfield, 2001b]), and indicate sulfur recycling.

We studied S cycling and its role for C metabolism in two peat soils at the upper and lower end of atmospheric S deposition in North America. The soils included the moss canopy, were contained in mesocosms, and subjected to water flow and sulfate deposition. Our specific objectives were (1) to examine BSR activity in the peat, (2) to seek evidence for a renewal of the dissolved sulfate pool along the flow path, and (3) to identify potential mechanisms by which the sulfate pool was maintained. BSR was inferred from concentration changes of sulfate and H$_2$S and enrichment of $^{34}$S-sulfate along the flow path (Figure 1) and incubation experiments in which BSR was chemically suppressed. Depth profiles of total reduced inorganic sulfur (TRIS) were examined in separately taken peat cores. Renewal of the sulfate pool was indirectly inferred from depth profiles of sulfate and H$_2$S concentrations and $^{34}$S contents in relevant S pools. The potential importance of DOM as an oxidant for reduced S was assessed from production rates of dissolved C in the mesocosms. We expected to find, regardless of soil origin and treatment, a retention of sulfate by the vegetation, low sulfate concen-

**Figure 1.** Schematic of some S transformations in saturated peat soils with vertical water flow, and expected shifts in $^{34}$S/$^{32}$S ratios associated with them. Sulfate passing the Sphagnum moss canopy may form organic sulfate esters (SO$_4^2$-R) or be reduced by sulfate reducing bacteria (SRB). Hydrogen sulfide may subsequently be chemically incorporated into reduced inorganic sulfur (TRIS) or carbon-bonded sulfur (CBS), or be chemically or microbially reoxidized to thiosulfate (S$_2$O$_3^{2-}$), for example by quinone moieties (DOM-Q) contained in dissolved organic matter that leaches from vegetation and the unsaturated zone. Sulfate reduction leaves the remaining sulfate enriched in $^{34}$S and the produced H$_2$S, TRIS, and CBS depleted in $^{34}$S along the flow path (strong isotope fractionation [Canfield, 2001a]). Replenishment of the sulfate pool from H$_2$S, TRIS, and CBS will reduce the $^{34}$S content of sulfate along the flow path (small isotope fractionation [Canfield, 2001b]), and indicate sulfur recycling.
of sulfate, representing a deposition of $S$ at $0.83$ and $3.33 \, g \, m^{-2} \, yr^{-1}$. Two mesocosms were exposed to high (MB1, ELA 1) and two to low (MB 2, ELA 2) levels of $S$ loading. The precipitation also contained $H_2O^+$ (92–358 $\mu$mol $L^{-1}$), $NO_3^-$ (40–120 $\mu$mol $L^{-1}$), $NH_4^+$ (40–120 $\mu$mol $L^{-1}$) and $Ca^{2+}$ (30 $\mu$mol $L^{-1}$), $Mg^{2+}$ (15 $\mu$mol $L^{-1}$), $Na^+$ (50 $\mu$mol $L^{-1}$), $K^+$ (5 $\mu$mol $L^{-1}$), and $Cl^-$ (150–265 $\mu$mol $L^{-1}$).

[11] Pore water (5 to 10 mL) was extracted from the suction samplers with stoppered syringes shortly before terminating the experiment and used to analyze dissolved inorganic carbon (DIC), DOC, $\mathrm{CH}_4$, sulfate, $\mathrm{H}_2\mathrm{S}$, $\mathrm{Fe}^{2+}$ and pH. The mesocosms were then transferred to the laboratory, closed on top by a cap and nitrogen was flushed through using a tube attached to the cap. After about 10 min, the mesocosms were drained through the horizontally inserted pore water samplers by gravity and $N_2$ pressure, while continuously flushing with $H_2$. Beginning with the uppermost soil, a depth increment of 6–12 cm was allowed to complete draining. Then the next depth increment further down was drained. This procedure minimized mixture of pore water from different depths. The pore water was frozen and stored at $-18^\circ C$ until analyzed. The peat core was transferred to a glove box that was subsequently filled with nitrogen. Then the cap was removed and the core extruded, dissected and peat sampled for incubation and solid phase analyses.

2.3. Analyses

2.3.1. Pore Water

[12] DOC was determined after filtration of extracted pore water with a syringe microfilter (0.45 $\mu$m, nylon) on a Shimadzu 5050 TOC analyzer. Dissolved inorganic carbon (DIC) and $\mathrm{CH}_4$ were determined on a Shimadzu Mini 2 gas chromatograph with methanizer in the gas phase of 1.8 mL vials, as described by Blodau et al. [2004]. Dissolved $O_2$ was determined amperometrically with a low-current electrode, and $\mathrm{H}_2\mathrm{~S}$ and pH potentiometrically ($\mathrm{Ag}/\mathrm{AgCl}$-glass-electrode, Watertest) on 0.5 to 3 mL of sample with a conventional meter (Orion). The sulfide electrode was conditioned with $\mathrm{H}_2\mathrm{~S}$ (100 $\mu$mol $L^{-1}$) prior to use. Dissolved $O_2$ contamination due to the sampling procedure was $\sim 0.5$ mg $L^{-1}$. Sulfate was determined by ion chromatography (Metrosep Anion Dual 1, at 0.5 mL min$^{-1}$ flow rate and chemical suppression). In a number of samples from the mesocosms, concentrations of formate, acetate, propionate, and butyrate were determined by HPLC at the end of the experiment. Dissolved Fe($II$) and Fe(total) were determined with the phenanthroline method and addition of ascorbic acid (Fe(total)) [Tamura et al., 1974].

2.3.2. Solid Phases

[13] We collected additional peat cores from the Mer Bleue peatland for the analysis of inorganic solid phase Fe and S. Reactive iron was determined after cold extraction for 24 hours with 1M $HCl$ [Wallmann et al., 1993], using the phenanthroline method. Iron(III) was calculated as the difference between Fe(tot) and Fe(II). The content of total inorganic reduced sulfur compounds (TRIS: $\mathrm{FeS}_2$, $\mathrm{FeS}$, $S^0$) was determined using the method of [Fossing and Jorgensen, 1989]. Frozen peat samples were freeze dried and the material boiled with $HCl$ (c = 5 $mol \, L^{-1}$) and $CrCl_2$...
was extracted from freeze dried sediment, 1985]. A preliminary calculation of BLODAU ET AL.: SULFUR CYCLE IN PEATS

Solution, followed by filtration of D2 in the low S treatment and /C176 (Figure 2). The majority of the whole sediment diffusion coefficient, cm s⁻¹/D2 m /C0 Cornel et al. by ebullition of CH by D2 solution (10%) was added to are required for the process in an elemental d m at 8 to 34 ¼ precipitate. The Ag cm s⁻¹/C176 volume of 5 mL of BaCl2 of the mesocosms ([300 mL per depth] obtained from the incremental drainage 2.4. Calculations

sulfate to complete the analysis. 34 S precipitate was dried, weighed, and stored for subsequent isotopic analysis. For analysis of 34 S/C2 ratios in total peat, the dry weight was dried at 70°C and finely ground. The sulfur isotope ratio of sulfate, total S, and TRIS was determined by converting either BaSO4, the peat sample, or Ag2S into SO2 in an elemental analyzer (EA 1500) coupled to a mass spectrometer (VG Prism II) in continuous flow mode. Sulfur isotope ratios are reported on the usual δ-scale in parts per thousand deviation relative to the internationally accepted standard V-CDT (troilite from the Canyon Diablo meteorite),

\[ \delta^{34}S_{[\text{‰}]} = \left( \frac{^{34}S/^{32}S_{\text{sample}}}{^{34}S/^{32}S_{\text{reference}}} - 1 \right) \times 1000. \]

3.1 Concentration Profiles and C Turnover

Potential production rates of DIC, DIC, CH4, and H2S were determined at four depths (~10, 24, 34, and 60 cm) under anaerobic conditions at the end of the experiment. Peat cores were extruded, dissected, and peat sampled under nitrogen in a glove chamber. The peat (50–100 g wet weight, 7.2 ± 3.2 g (s.d.) dry weight) was placed in 125–250 mL, rubber-stoppered Erlenmeyer flasks and fully immersed in daerated water. All flasks were filled with solution containing 500 μmol L⁻¹ of Na2SO4. One half of the flasks was prepared by adding 1000 μmol L⁻¹ of (NH4)2MoO4, a specific inhibitor of BSR [Oremland and Capone, 1988]. The water was sampled through suction samplers inserted though the rubber stoppers that closed the flasks. Within 7 to 9 days, ~10 mL of water were extracted on five to six occasions and DOC, DIC, CH4, H2S, Fe²⁺ concentrations and pH determined. The water was replaced with oxygen-free solution. Rates were determined from volume-corrected linear regression of concentration over time.

3. Results

3.1 Concentration Profiles and C Turnover

Concentrations of dissolved sulfate in the irrigation water were 26 μmol L⁻¹ in the low S treatment and 104 μmol L⁻¹ in the high S treatment, which would increase to 47 and 191 μmol L⁻¹ after correction for evaporation. Sulfate concentrations below the water table were low, ranging from 10 to 20 μmol L⁻¹ (Figure 2). The majority...
of the deposited sulfur was thus retained by the Sphagnum canopy, capillary fringe, and at the water table. Oxygen concentrations decreased within three centimeters below the water table to background contamination levels <15 μmol L⁻¹ (Figure 2). Hydrogen sulfide concentrations ranged from non-detectable to 9 μmol L⁻¹ in the MB 1 mesocosm (Figure 2). Maximum H₂S concentrations were detected at depths of 10 to 30 cm, depending on the mesocosm. Ferrous iron concentrations ranged from <10 to 50 μmol L⁻¹. In the MB mesocosms, Fe²⁺ was mobilized within the first 10 cm below the water table and immobilized deeper. Concentrations of Fe²⁺ were less variable in the ELA mesocosms.

Figure 2. Depth profiles of dissolved oxygen, sulfate, ferrous iron, and hydrogen sulfide in mesocosms. The uppermost value indicates the position of the water table. Note different concentration scale for hydrogen sulfide (bottom).

Dissolved C was dominated by the DOC fraction, with concentrations ranging from 1.4 to 9 mmol L⁻¹ (Figure 3) and with the percolate from the Sphagnum canopy and thin unsaturated zone containing 1.4 to 3.5 mmol DOC L⁻¹. In the MB mesocosms, DOC concentrations continued to increase with depth, reaching a maximum at 30 to 40 cm, whereas this increase was small in the ELA mesocosms. DIC concentrations increased from 0.12–0.35 mmol L⁻¹ at the water table to 1.7–2.3 mmol L⁻¹. Methane concentrations were much lower than DIC concentrations and increased from 0.009–0.08 mmol L⁻¹ at the water table to 0.53–0.63 mmol L⁻¹ at 20 to 60 cm. Acetate, propionate, and butyrate remained below the detection limit of ~0.05 mmol L⁻¹ in the subset of samples analyzed. Maximum production of DIC was 21 to 93 nmol cm⁻³ d⁻¹ and of CH₄ 6.2 to 11.5 mmol cm⁻³ d⁻¹. The fastest rates generally occurred in the surface peat for DIC and at depths of 15 to 30 cm for CH₄ (Table 1). The ELA 2 mesocosm was an exception, with the peat below 10 cm being fairly inactive. Standardized to the peat surface, the magnitude of rates was similar among cores and decreased in the order DOC > DIC > CH₄ (Table 2). Production of DOC, however, was considerably faster in the MB soil.

3.2. Solid Phases

In the cores extracted separately from the Mer Bleue peatland, TRIS occurred in all depth layers, even those (<20 cm) typically above the water table (Figure 4), indicating in situ formation of iron sulfides. A maximum of 10 to 19 μmol g⁻¹ of TRIS was found at depths of 25 to 35 cm where elemental S reached a maximum of 1 to 2 μmol g⁻¹. Reactive ferric and ferrous iron concentration peaked in the depth layer of 15 cm at low concentrations of 15 to 25 μmol g⁻¹.

3.3. Sulfur Isotopes

The δ³⁴S values of total sulfur decreased with depth in all cores, from 5.0–6.0‰ to 1.9–2.4‰ in ELA and 5.7–6.8‰ to 4.0–4.3‰ in MB mesocosms (Figure 5). In the lower layers of the MB cores, however, δ³⁴S values of total S increased. Sulfur isotope ratios in TRIS followed roughly the same trend, but the differences in δ³⁴S with depth were more pronounced. In MB 1, δ³⁴S values distinctly decreased from 7.2‰ (10 cm) to 1.4‰ (48 cm). The δ³⁴S value of sulfate in the irrigation water was 7.4‰ and 0.2 and 2.2‰ smaller just below the water table. In all four cores, sulfate was enriched in ³⁴S at depths of 10 to 25 cm, where δ³⁴S values reached 10.1‰ (MB 1 and 2), 11.8‰ (ELA 2) and 13.4‰ (ELA 1). Below that maximum, δ³⁴S values decreased but mostly remained well above values of total S and TRIS.

Figure 3. Depth profiles of dissolved inorganic carbon (DIC), CH₄, and dissolved organic carbon (DOC) in mesocosms. Note different concentration scale for DOC (bottom).
3.4. Incubation Experiments

[25] Dissolved carbon and H$_2$S concentrations increased steadily in most incubation flasks (see example in Figure 6), sometimes after an initial lag time (H$_2$S and CH$_4$). Ferric iron and proton concentrations mostly decreased or were without trends. The anaerobic potential DIC production rates mostly decreased with depth (Figure 7). Addition of MoO$_4^{2-}$ decreased averaged DIC production rates by 20 to 50% (paired t-test, *P < 0.05; Figure 7 and Table 3). The potential CH$_4$ production rates peaked on average at 24 cm and 0.75 ± 0.32 μmol g$^{-1}$ d$^{-1}$ (Figure 7). Addition of MoO$_4^{2-}$ decreased potential CH$_4$ production rates by a factor of 2–8 times (paired t-test, *P < 0.05) at depths of 10 and 24 cm but not at greater depths (Table 3 and Figure 7). Potential rates were on average larger than the corresponding diffusive-advective in situ rates by a factor 2 (10 cm) to 26 (60 cm). Potential DOC production rates decreased on average from 17 to 4 μmol g$^{-1}$ d$^{-1}$ with depth. Addition of MoO$_4^{2-}$ resulted in increased rates (factor 1.5 to 2.1) at depths of 10–24 cm, compared to the controls.

[24] Potential H$_2$S production rates ranged from 0.007 to 2.6 μmol g$^{-1}$ d$^{-1}$ (Figure 7). Production was fastest in the mesocosm and the layers that were characterized by the highest in situ concentrations of H$_2$S (MB1, 10 to 25 cm; Figure 2) and a similar pattern was found in ELA 1. In the surface layers of MB 2 and ELA 2, however, H$_2$S was also produced, although we were unable to detect the gas in situ. A statistical correlation between the variables could thus not be established. Addition of molybdate suppressed H$_2$S production completely.

4. Discussion

[25] Bacterial sulfate reduction in northern peatlands has been investigated in several studies using $^{35}$S-radiotracer incubation techniques. On the basis of the incorporation of the tracer in iron sulfides and elemental S, BSR rates of 2.5 to 180 nmol cm$^{-3}$ d$^{-1}$ have been reported from peats at temperatures of 20 to 25°C and across the gradient of S deposition of northern peatlands [Vile et al., 2003b]. Given the incorporation of $^{35}$S in organic matter, BSR rates can be substantially higher [Vile et al., 2003a]. Sulfate reduction rates in peat bogs may overlap with rates in environments where sulfate is not limiting BSR, such as salt marshes [Howarth and Jorgensen, 1984], sediments of lakes polluted by acid mine drainage [Blodau et al., 1998], and marine sediments [Canfield et al., 1993]. On the basis of the small sulfate pools present, which usually range from 10 to 300 μmol L$^{-1}$ in peat bogs, a turnover time of sulfate on the order of hours to weeks has been reported [Nedwell and Watson, 1995; Wieder and Lang, 1988]. This has led to the suggestion of an anaerobic recycling of sulfur to oxidized forms and subsequent repeated reduction [Nedwell and Watson, 1995; Wieder and Lang, 1988; Vile et al., 2003a]. It was our objective to find evidence for such a recycling process.

[26] The mesocosms were exposed to constant water tables for more than 200 days, in which the initial pore water had been mostly replaced by irrigation water [Blodau and Moore, 2002]. Anaerobic conditions and adjustment to input of sulfur was thus established for long periods of time before sampling. A reoxidation of reduced sulfur by water table fluctuations, to which a renewal of sulfate pools in peats has been frequently attributed [Bayley et al., 1986; Lazerte, 1992], can thus be discarded.

[27] Most of the deposited sulfur did not reach the anaerobic zone below the water table. Between 76 and 94% of the deposited sulfate was retained by the Sphagnum layer and unsaturated zone in the mesocosms when the water table was near the surface, as was previously reported by Moore et al. [2005]. The influence of long-term in situ loading with sulfur (MB high and ELA low) and the short-term sulfate loading in the experiments (MB1 and ELA 1 high, and MB2 and ELA2 low) thus seemed to be of minor importance for sulfate supply to BSR. Because of sulfate retention, sulfate concentrations typically ranged from 5 to 20 μmol L$^{-1}$ below the water table, long considered the lower boundary for dissipilatory activity of sulfate reducers [Ingversen et al., 1981, 1984; Lovley and Klug, 1983]. Still, sulfate was reduced in our peat mesocosms. Hydrogen sulfide was present in the pore waters of three of the four mesocosms, and TRIS accumulated at depths of 10 to 30 cm in situ at the Mer Bleue site. Evidence for BSR was further provided by the enrichment of $^{33}$S in the sulfate pool at these depths in all cores (Figure 5). Using natural populations of SRB, sulfate reduction has resulted in isotopic fractionation with enrichment factors typically between 8% and 40%, depending on temperature, substrate availability, and specific sulfate reduction rates [Canfield, 2001a]. Our results thus suggest that BSR remained an

| Table 1. Net Production of DIC, CH$_4$, and DOC in Depth Segments of Peat Mesocosms
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*aNet production is in nmol cm$^{-3}$ d$^{-1}$ and depth segments are in centimeters.

| Table 2. Depth-Integrated Production in Mesocosms
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*aNumbers in depths are in centimeters. Production is mmol m$^{-3}$ d$^{-1}$. 
active process even at very low sulfate concentrations, corroborating results by Nedwell and Watson [1995] from an ombrotrophic bog in Britain. The results further support the idea of sulfate reducers accounting for a substantial fraction of the anaerobic metabolic activity. We observed DIC production that could not be attributed to methanogenesis or the utilization of nitrate and ferrous iron. Nitrate concentrations were <5 µmol L\(^{-1}\) [Blodau et al., 2006; Moore et al., 2005]. Concentrations of ferrous iron, which mainly functioned as a scavenger for sulfide, remained low (<50 µmol L\(^{-1}\)) in the pore waters and solid phase. Incubation experiments with peat from the Mer Bleue site had shown only small rates (<0.1 µmol g\(^{-1}\) d\(^{-1}\)) of ferrous iron release, independent of sulfate and acetate availability in the slurries, and may have accounted for 0.3–0.4% of DIC production [Blodau et al., 2002]. In contrast, addition of molybdate suppressed \(\text{H}_2\text{S}\) formation fully and DIC production by 20 to 50% in the experiments reported here. Molybdate blocks the activation of sulfate by ATP through formation of an unstable APMo analogue to APS within sulfate respiring bacteria [Oremland and Capone, 1988]. As sulfate respiring bacteria must first expend energy in form of ATP before generating energy from respiration, ATP pools are rapidly depleted in presence of molybdate, thereby likely causing cell death [Taylor and Oremland, 1979]. Such a process and associated cell lysis might have contributed to the increased net DOC production (Table 3), apart from physicochemical effects and the decrease in DIC production that occurred. Interestingly, suppression of BSR by addition of molybdate was accompanied by lower methane production rates (Table 3), suggesting that SRB and methanogens did not only compete for substrates. In agreement with this interpretation, also the addition of sulfate partly had little effect on methane production in incubation experiments with peats from the Mer Bleue bog [Blodau and Moore, 2003]. Both findings are not implausible assuming that a fraction of SRB used lactate as substrate and produced acetate, which is an important substrate for methanogenesis in peatlands [Avery et al., 2003; Hornibrook et al., 1997].

DIC release unexplained by \(\text{CH}_4\) production amounted to 0–81 nmol cm\(^{-2}\) d\(^{-1}\). In terms of electron

Figure 4. Depth profiles of total reduced inorganic sulfur (TRIS), elemental S, and HCl-dissolvable ferric and ferrous iron in cores sampled separately at the Mer Bleue site in May 2003.

Figure 5. Isotopic signatures (\(\delta^{34}\text{S}\)) of sulfate, total S, and TRIS, as obtained from drained pore water (sulfate) and solid phase extraction.

Figure 6. Time course of an exemplary incubation showing a high potential hydrogen sulfide production (MB1, depth 24 cm).
equivalents, unexplained CO$_2$ production rates were similar to BSR rates (7–38 nmol cm$^{-3}$ d$^{-1}$) in a Cedar Swamp [Spratt and Morgan, 1990] and northern peatlands [Vile et al., 2003a, 2003b]. The calculated differences between net DIC and CH$_4$ production were probably underestimated as DIC is commonly reduced to CH$_4$ in peatlands [Avery et al., 2003; Hornibrook et al., 1997]. This holds true at least in the upper peat layers, where CH$_4$ concentrations were <358–389 μmol L$^{-1}$ and thus insufficient to allow for ebullition, a transport mechanism that we could not account for in the dissolved C mass balance approach.

The occurrence and potential significance of BSR in the mesocosms raise the question whether and where sulfur recycling occurred and what processes may have been involved. Instructive in this respect is an examination of the $\delta^{34}$S profiles (Figure 5). The $\delta^{34}$S-sulfate values decreased after a maximum at a depth of 10 to 20 cm and roughly followed the $\delta^{34}$S profiles of TRIS and total sulfur, with the exception of the deepest sample in ELA 1 (Figure 5). This pattern contradicts a Rayleigh distillation along a flow or diffusion path in a semi-closed system, which will result in progressively increasing $\delta^{34}$S values with depth for sulfate, and a growing difference between $\delta^{34}$S values of sulfate and reduction products, such as H$_2$S and TRIS [Jorgensen, 1979]. The decrease in $\delta^{34}$S-sulfate values with depth thus requires replenishment of the sulfate pool from the peat matrix. Taking into account that the sulfate concentration profiles did not show a net mass loss in the zone of 10–20 cm (Figure 5), where $^{34}$S-enriched sulfate accumulated and sulfate was reduced, recycling must already have occurred in the zone of active BSR near the surface.

[31] The processes involved in the recycling cannot be unequivocally identified on the basis of our results but some insight can again be gained from combining concentration profiles, isotopes and incubations. Heitmann and Blodau [2006] have previously reported that peat DOM can reoxidize H$_2$S chemically to thiosulfate which can be used by SRB and thus support their dissimilatory activity and growth [Jorgensen, 1990; Jorgensen and Bak, 1991; Widdel, 1988]. Thiosulfate can further be disproportionated into H$_2$S and sulfate by microbial mediation, thus resupplying the sulfate pool [Habicht et al., 1998]. Although we cannot directly support such a mechanism by thiosulfate measurements and changes in electron accepting capacities of the DOM with depth, it seems likely that this process occurred in the uppermost peat layer, where DOM must have been leached in its oxidized form. There we determined the highest potential H$_2$S production rates (Figure 7) and observed the largest suppression of DIC production (Table 1 and Figure 7). Hydrogen sulfide was furthermore net consumed by some process as indicated by a decreasing concentration below a peak of maximum concentration (Figure 2).

Figure 7. Potential production rates of DIC, CH$_4$, and H$_2$S. In all incubations sulfate (0.5 mM) and in one set ammonium molybdate (1 mM) was added to the incubation water. Note different scale for DIC production.

Table 3. Change in Production With Molybdate Addition in Peat Incubations With MB Peat$^a$

<table>
<thead>
<tr>
<th>Depth, cm</th>
<th>DIC</th>
<th>CH$_4$</th>
<th>DOC</th>
</tr>
</thead>
<tbody>
<tr>
<td>10</td>
<td>$-50 \pm 11$</td>
<td>$-37 \pm 92$</td>
<td>$+74 \pm 4$</td>
</tr>
<tr>
<td>24</td>
<td>$-22 \pm 28$</td>
<td>$-47 \pm 19$</td>
<td>$+98 \pm 82$</td>
</tr>
<tr>
<td>36</td>
<td>$-42 \pm 11$</td>
<td>$-42 \pm 36$</td>
<td>$+19 \pm 47$</td>
</tr>
<tr>
<td>60</td>
<td>$-29 \pm 13$</td>
<td>$-7 \pm 20$</td>
<td>$+27 \pm 3$</td>
</tr>
</tbody>
</table>

$^a$Change is given as percent ± s.d. Here n = 3.
Mass balance considerations suggest that production of electron accepting capacity related to DOM alone could not account for the DIC production, even if a maximum of electron transfer capacities was assumed. The electron accepting capacity of fully oxidized Mer Bleue DOM toward H$_2$S may reach 6 mmol g$^{-1}$ C, or one electron transfer for 14 C atoms [Heitmann et al., 2007]. Using this value, only 1.5 to 5% of DIC production unexplained by methane production can be accounted for in the uppermost layer. In deeper layers, this percentage was higher but it is not clear in what redox state DOM was released. This consideration leaves the capacity of the solid phase organic matter to reoxidize H$_2$S unaccounted for. Information about this capacity is not available yet and its contribution to respiration cannot be evaluated.

A complete recycling of H$_2$S back to sulfate at depths of 10 to 20 cm would not produce a peak in $^{34}$S-enriched sulfate as observed but would leave $^{34}$S values unchanged. Part of the reduced sulfate must have left the reoxidation cycle in the uppermost layers, leading to a net enrichment in $^{34}$S (Figure 1). H$_2$S can be incorporated into organic matter [Brown, 1986; Heitmann and Blodau, 2006], for example by Michael’s addition reaction with quinone moieties [Perlinger et al., 2002]. Formation of CBS may often, but not always, predominantly follow the production of H$_2$S by BSR in peat soils [Spratt and Morgan, 1990; Vile et al., 2003a; Wieder and Lang, 1988]. Incorporation of H$_2$S into DOM also occurred in pore water from the Mer Bleue peatland (T. Heitmann and C. Blodau, unpublished data, 2005).

At greater depths, sulfate was depleted in $^{34}$S and the sulfate pool must have been replenished either by a reoxidation process providing sulfate from hydrogen sulfide, or TRIS, or from organic sulfur, which had a much lower $^{34}$S signal than sulfate (Figures 1 and 5). A release of S deeper into Sphagnum peat has been inferred earlier on the basis of increasing $^{34}$S values of peat with depth [Novák et al., 1994]. Among the possible mechanisms for S release, hydrolysis of organic ester-bound sulfate is probably common in peatlands [Mandernack et al., 2000]. It does not seem to be a viable explanation for the decreasing $^{34}$S values of sulfate, however, because at greater depths ester-bound sulfate should have been enriched in $^{34}$S itself. It seems likely that at least part of the sulfate-sulfur stemmed from the TRIS fraction, because TRIS contents decreased below a maximum at 20 to 30 cm in peat cores from Mer Bleue (Figure 4), similarly as in other Sphagnum bogs [Novák et al., 1994]. With burial, TRIS contents should increase owing to selective accumulation with C loss over time. This was not the case and suggests that TRIS was labile and reoxidized, as previously documented experimentally by Wieder and Lang [1988]. The same conclusion applies to elemental S, which is an important intermediate in the microbial reoxidation of hydrogen sulfide and can be both reduced and oxidized [Canfield and Thamdrup, 1996].

In previous studies, the deposition of S has had a larger effect on CH$_4$ production and emissions than expected, on the basis of sulfate and CH$_4$ mass balances [Dise and Verry, 2001; Gauci et al., 2002]. A continuous reoxidation of reduced sulfur to sulfate under anaerobic conditions should contribute to this large effect and is likely based on previous studies and our results. The addition of molybdate to peat furthermore showed that not much dissolved sulfate was needed to maintain a high activity of SRB. This is also in agreement with earlier work on this issue [Vile et al., 2003a]. Very low sulfate concentration levels of 5 to 20 $\mu$mol L$^{-1}$ were attained after equilibration of the mesocosms to the chosen conditions of water table and S deposition. Hence the activity of SRB in our peat soils must have been constrained by the kinetics of S oxidation or by the provision of electron accepting capacity that fuelled the oxidation. If this interpretation is correct, and if the rate of S oxidation is proportional to the size of reduced S pools, an increase in sulfate pool sizes by sulfate deposition may result in increased BSR rates, although the sulfate pool itself cannot maintain BSR for extended periods. We cannot refute or confirm this conclusion on the basis of the data obtained. It is in agreement with the observation that rates of BSR increase and methane production decrease with S loading, as documented by Vile et al. [2003b] and with the preferential retention and accumulation of atmospherically deposited S in ombrotrophic peatlands [Moore et al., 2005].

5. Conclusion

We examined sulfur cycling in intact peat mesocosms from two peatlands at natural background (ELA) and moderately elevated sulfate deposition (MB), and two levels of experimental sulfate deposition. Sulfate was reduced at sulfate concentrations <20 $\mu$mol L$^{-1}$ and SRB were important for the anaerobic electron flow, as their inhibition substantially lowered potential CO$_2$ production rates. We suggest that BSR was partly maintained by a recycling of reduced inorganic sulfur that was “leaky” with respect to organic S. In the upper peat layers, $^{34}$S-enriched sulfate accumulated in the absence of net sulfate loss from pore water. This suggested a mode of S cycling in which sulfate was reduced and the sulfide produced only partly reoxidized, and partly incorporated into organic matter and TRIS. Deeper in the peat, release and reoxidation of S stemming from organic matter or TRIS became more important, as indicated by decreasing $^{34}$S values in sulfate. The results of this study thus support earlier investigations that documented high rates of BSR in presence of small sulfate pools and provide support for a reoxidation of reduced S in peats poor in inorganic oxidants.

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